

Chapter 8 Ionic Compounds Answer Key

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Chapter 8 Ionic Compounds Answer

solutions of electrolytes in Chapter 15. 218 Chapter 8 Ionic Compounds Figure 8-7 Aragonite (CaCO_3), barite (BaSO_4), and beryl ($\text{Be}_3\text{Al}_2\text{Si}_6\text{O}_{18}$) are examples of minerals that are ionic compounds. The ions that form them are bonded together in a crystal lattice. Melting and Boiling Points of Some Ionic Compounds

Compound	Melting point ($^{\circ}\text{C}$)	Boiling point ($^{\circ}\text{C}$)
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8. In a crystal lattice of an ionic compound, a. ions of a given charge are clustered together, far from ions of the opposite charge. b. ions are surrounded by ions of the opposite charge. c. a sea of electrons surrounds the ions. d. neutral molecules are present. Chemistry: Matter and Change Chapter 8 Study Guide for Content Mastery 44

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Predict the empirical formulas of the ionic compounds ...

How Do Ionic Compounds Form? Ionic bonds form because positive ions are attracted to negative ions. When ionic bonds form, the number of electrons lost by metal atoms equals the number of electrons gained by nonmetal atoms. The ions that form an ionic compound are charged, but the compound they form is neutral. That is because

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SECTION 2 Ionic Bonds

Naming Ions. The name of a monatomic cation is simply the name of the element followed by the word ion. Thus, Na^+ is the sodium ion, Al^{3+} is the aluminum ion, Ca^{2+} is the calcium ion, and so forth. We have seen that some elements lose different numbers of electrons, producing ions of different charges (Figure 3.3).

5.7: Naming Ionic Compounds - Chemistry LibreTexts

Recognizing Ionic Compounds. There are two ways to recognize ionic compounds. First, compounds between metal and nonmetal elements are usually ionic. For example, CaBr_2 contains a metallic element (calcium, a group 2 (or 2A) metal) and a nonmetallic element (bromine, a group 17 (or 7A) nonmetal). Therefore, it is most likely an ionic compound.

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