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from the loss or gain of one or more valence electrons, respectively. Key 1. When the metals in Group 1A, 2A, and 3A lose electrons, they form cations with positive charges equal to their group number.

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Section 9.4 - Naming and Writing Formulas for Acids and Bases. An acid is

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a compound that produces H^+ ions when it dissolves in water. The formula for an acid normally starts with an H. When naming acids, you should first determine whether or not the acid contains oxygen.

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COMPOUNDS (pages 268–270) This
section explains the rules for naming

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CHAPTER 9 Assessment 42. a. 2- b. 1+
c. 1- d. 3+ 43. a. 2+ b. 2+ c. 3+ d.1+

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44. a. barium ion b. iodide ion c. silver ion d. mercury(II) ion 45. cyanide, CN^- and hydroxide, OH^- 46. a. hydroxide ion b. lead(IV) ion c. sulfate ion d. oxide ion 47. zero; A compound is electrically neutral. 48. The symbols for the cation and

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Chapter 9 "Chemical Names and Formulas" Tools. Copy this to my account; E-mail to a friend; Find other activities; ... In samples of any chemical compound, the masses of the elements are always in the same proportions any atom or group of atoms that has a negative charge ... place cation name first followed by the anion name: naming

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polyatomic ...

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Chapter 9 - Chemical Names and Formulas - Mrs. Marquart's ...

Chapter 9 Chemical Names and
Formulas 221 Name Date 4 Write
formulas for the following ionic

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compounds a sodium phosphate c
sodium hydroxide e ammonium chloride
b magnesium d potassium cyanide 5
Write formulas for compounds formed
from these pairs of ions b , 6 Name the
following compounds ...

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Chapter 9 “Chemical Names and Formulas” ... in parenthesis after the name of the metal (Table 9.2, p.255)
Predicting Ionic Charges Some of the post-transition elements also have more than one possible oxidation state. Tin (II) = Sn^{2+} Lead (II) = Pb^{2+} Tin (IV) = Sn^{4+} Lead (IV) = Pb^{4+}

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Chapter 9 Chemical Names and Formulas 211. Section Review.

Objectives.

- Determine the charges of monatomic ions by using the periodic table and write the names of the ions.
- Define a polyatomic ion and write the names and formulas of the most common polyatomic ions.
- Identify the

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two common endings for the names of most polyatomic ions.

05 CTR ch09 7/9/04 3:29 PM Page 211 NAMING IONS 9

Chapter 9 Chemical Names and Formulas 221 Name Date 4 Write formulas for the following ionic compounds a sodium phosphate c

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sodium hydroxide e ammonium chloride
b magnesium d potassium cyanide 5
Write formulas for compounds formed
from these pairs of ions b , 6 Name the
following compounds ...

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