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To describe electrons at the nanoscale, we will begin with an introduction to the principles of quantum mechanics, including quantization, the wave-particle duality, wavefunctions and Schrödinger's equation.

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1 Problems Chapter 1: Introduction to Nanoelectronics 2
Problems Chapter 2: Classical Particles, Classical Waves, and Quantum Particles 2.1. What is the energy (in Joules and eV) of a photon having wavelength 650 nm? Repeat for an electron having the same wavelength and only kinetic energy. Solution: For the photon, $p = hc/E$; $E = hc/\lambda = hc/650 \times 10^{-9} \dots$

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